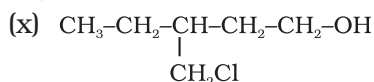
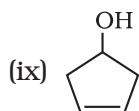
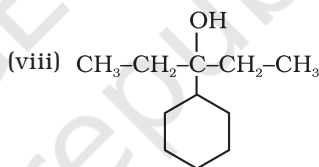
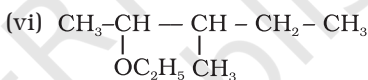
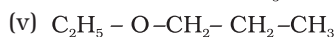
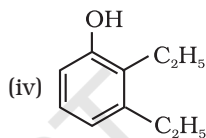
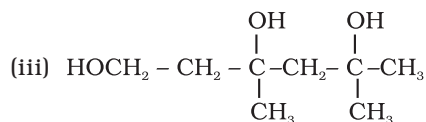
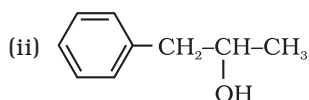
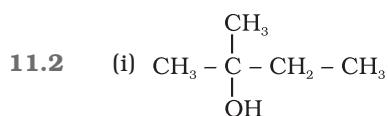


Answers to Some Questions in Exercises

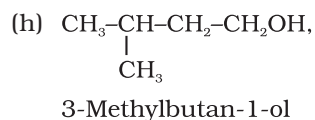
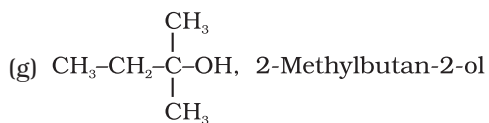
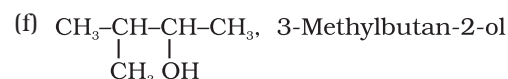
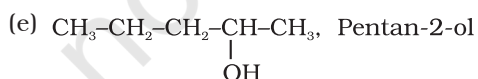
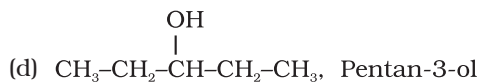
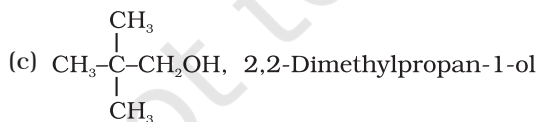
UNIT 11

- 11.1** (i) 2,2,4-Trimethylpentan-3-ol
 (iii) Butane-2,3-diol
 (v) 2-Methylphenol
 (vii) 2,5-Dimethylphenol
 (ix) 1-Methoxy-2-methylpropane
 (xi) 1-phenoxyheptane

- (ii) 5-Ethylheptane-2,4-diol
 (iv) Propane-1,2,3-triol
 (vi) 4-Methylphenol
 (viii) 2,6-Dimethylphenol
 (x) Ethoxybenzene
 (xii) 2-Ethoxybutane

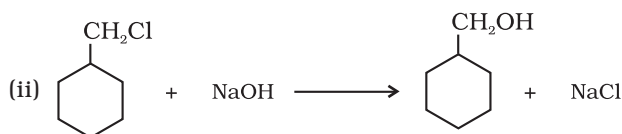
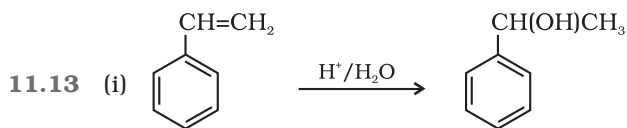


- 11.3 (i)** (a) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$, Pentan-1-ol;
 (b) $\text{CH}_3 - \text{CH}_2 - \underset{\text{CH}_3}{\text{CH}} - \text{CH}_2 - \text{OH}$, 2-Methylbutan-1-ol;



11.4 Hydrogen bonding in propanol.

- 11.5** Hydrogen bonding between alcohol and water molecules.
11.8 o-Nitrophenol is steam volatile because of intramolecular hydrogen bonding.
11.12 **Hint:** Carryout sulphonation followed by nucleophilic substitution.

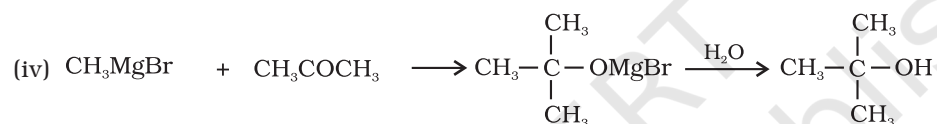
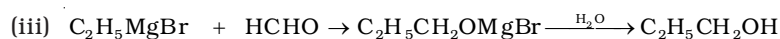


11.14 Reaction with (i) sodium and (ii) sodium hydroxide

11.15 Due to electron withdrawing effect of nitro group and electron releasing effect of methoxy group.

11.20 (i) Hydration of Propene.

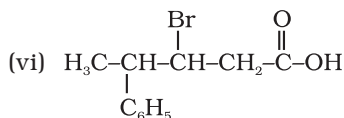
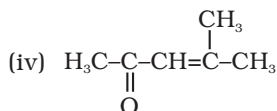
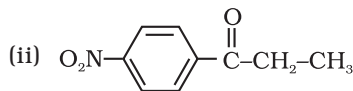
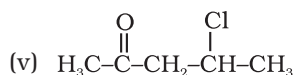
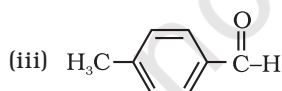
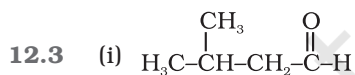
(ii) By nucleophilic substitution of $-\text{Cl}$ in benzyl chloride using dilute NaOH .

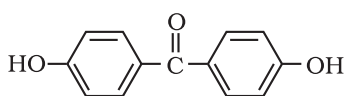


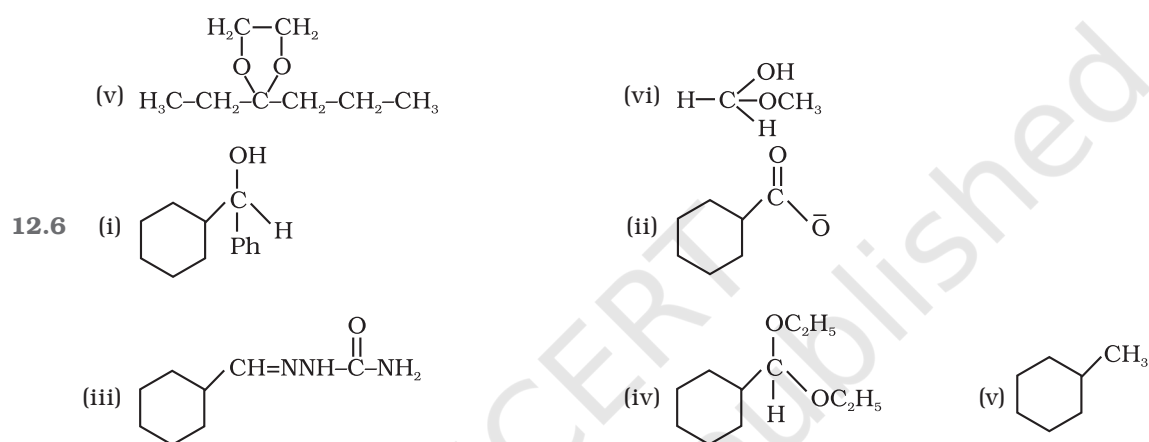
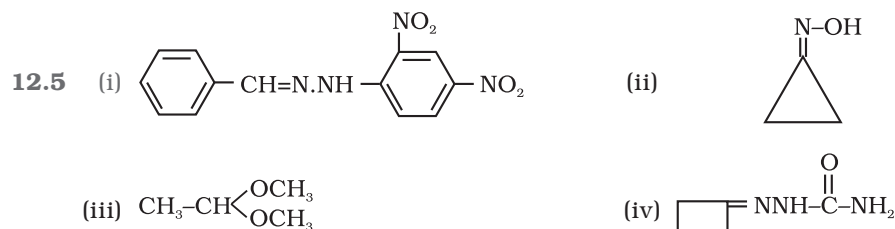
- 11.23** (i) 1-Ethoxy-2-methylpropane.
 (ii) 2-Chloro-1-methoxyethane.
 (iii) 4-Nitroanisole.
 (iv) 1-Methoxypropane.
 (v) 1-Ethoxy-4,4-dimethylcyclohexane.
 (vi) Ethoxybenzene.

UNIT 12

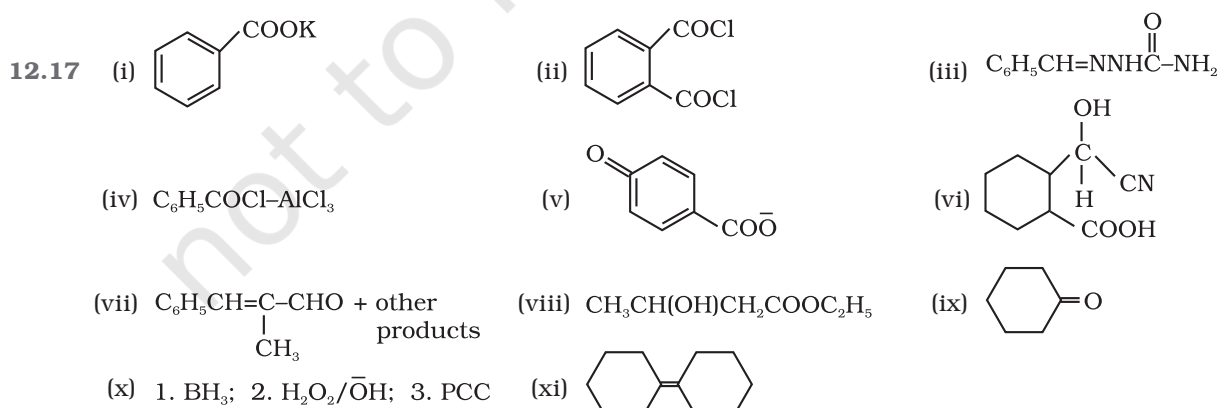
- 12.2** (i) 4-Methylpentanal (ii) 6-Chloro-4-ethylhexan-3-one
 (iii) But-2-enal (iv) Pentane-2,4-dione
 (v) 3,3,5-Trimethylhexan-2-one (vi) 3,3-Dimethylbutanoic acid
 (vii) Benzene -1,4-dicarbaldehyde



- (vii)  (viii) $\text{H}_3\text{C}-\text{C}\equiv\text{C}-\text{CH}=\text{CH}-\overset{\text{O}}{\parallel}{\text{C}}-\text{OH}$
- 12.4** (i) Heptan-2-one (ii) 4-Bromo-2-methylhexanal (iii) Heptanal
 (iv) 3-Phenylprop-2-enal (v) Cyclopentanecarbaldehyde (vi) Diphenylmethanone



- 12.7** (ii), (v), (vi), (vii): Aldol condensation. (i), (iii), (ix) Cannizzaro reaction. (iv), (viii) Neither.
- 12.10** 2-Ethylbenzaldehyde (draw the structure yourself).
- 12.11** (A) $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOCH}_2\text{CH}_2\text{CH}_2\text{CH}_3$, butyl butanoate.
 (B) $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$ (C) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$. Write equation yourself.
- 12.12** (i) Di-tert-butyl ketone < Methyl tert-butyl ketone < Acetone < Acetaldehyde
 (ii) $(\text{CH}_3)_2\text{CHCOOH}$ < $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$ < $\text{CH}_3\text{CH}(\text{Br})\text{CH}_2\text{COOH}$ < $\text{CH}_3\text{CH}_2\text{CH}(\text{Br})\text{COOH}$
 (iii) 4-Methoxybenzoic acid < Benzoic acid < 4-Nitrobenzoic acid < 3,4-Dinitrobenzoic acid.



- 12.19** The compound is methyl ketone and its structure would be: $\text{CH}_3\text{COCH}_2\text{CH}_2\text{CH}_3$

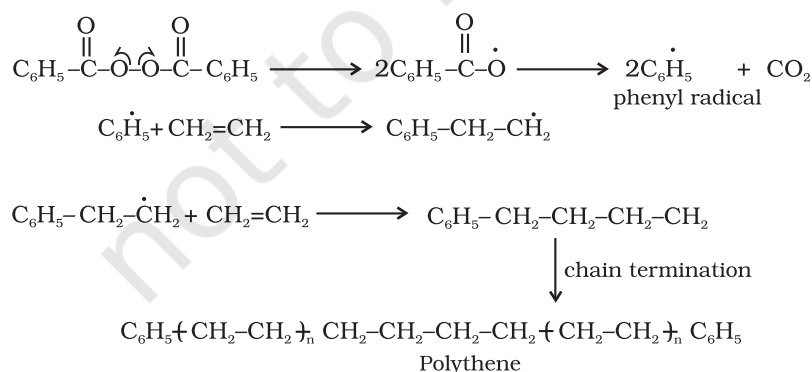
UNIT 13

- 13.1** (i) 1-methylethylamine or propan-2-amine (ii) Propan-1-amine
 (iii) N-methyl-2-methylethylamine or N-methylpropan-2-amine (iv) 2-methylpropan-2-amine
 (v) N-methylbenzenamine or N-methylaniline (vi) N-Ethyl-N-methylethanamine
 (vii) 3-Bromoaniline or 3-Bromobenzenamine
- 13.4** (i) $\text{C}_6\text{H}_5\text{NH}_2 < \text{C}_6\text{H}_5\text{NHCH}_3 < \text{C}_2\text{H}_5\text{NH}_2 < (\text{C}_2\text{H}_5)_2\text{NH}$
 (ii) $\text{C}_6\text{H}_5\text{NH}_2 < \text{C}_6\text{H}_5\text{N}(\text{CH}_3)_2 < \text{CH}_3\text{NH}_2 < (\text{C}_2\text{H}_5)_2\text{NH}$
 (iii) (a) p-nitroaniline < aniline < p-toluidine
 (b) $\text{C}_6\text{H}_5\text{NH}_2 < \text{C}_6\text{H}_5\text{NHCH}_3 < \text{C}_6\text{H}_5\text{CH}_2\text{NH}_2$
 (iv) $(\text{C}_2\text{H}_5)_3\text{N} > (\text{C}_2\text{H}_5)_2\text{NH} > \text{C}_2\text{H}_5\text{NH}_2 > \text{NH}_3$ (v) $(\text{CH}_3)_2\text{NH} < \text{C}_2\text{H}_5\text{NH}_2 < \text{C}_2\text{H}_5\text{OH}$
 (vi) $\text{C}_6\text{H}_5\text{NH}_2 < (\text{C}_2\text{H}_5)_2\text{NH} < \text{C}_2\text{H}_5\text{NH}_2$

UNIT 15

- 15.1** Polymer is a high molecular mass macromolecule consisting of repeating structural units derived from monomers.
 Monomer is a simple molecule capable of undergoing polymerisation and leading to the formation of the corresponding polymer.
- 15.2** Natural polymers are high molecular mass macromolecules and are found in plants and animals. The examples are proteins and nucleic acids.
 Synthetic polymers are man-made high molecular mass macromolecules. These include synthetic plastics, fibres and rubbers. The two specific examples are polythene and dacron.
- 15.4** Functionality is the number of bonding sites in a monomer.
- 15.5** Polymerisation is a process of formation of a high molecular mass polymer from one or more monomers by linking together of repeating structural units with covalent bonds.
- 15.6** Since the unit $-(\text{NH}-\text{CHR}-\text{CO})_n-$ is obtained from a single monomer unit, it is a homopolymer.
- 15.7** On the basis of molecular forces present between the chains of various polymers, the classification of polymers is given as follows.
 (a) Elastomers (b) Fibres (c) Thermoplastics and (d) Thermosetting plastics.
- 15.8** In addition polymerisation, the molecules of the same or different monomers add together to form a large polymer molecule. Condensation polymerisation is a process in which two or more bi-functional molecules undergo a series of condensation reactions with the elimination of some simple molecules and leading to the formation of polymers.
- 15.9** Copolymerisation is a process in which a mixture of more than one monomeric species is allowed to polymerise. The copolymer contains multiple units of each monomer in the chain. The examples are copolymers of 1,3-butadiene and styrene and 1, 3-butadiene and acrylonitrile.

15.10



- 15.11** A thermoplastic polymer can be repeatedly softened on heating and hardened on cooling, hence it can be used again and again. The examples are polythene, polypropylene, etc.
 A thermosetting polymer is a permanent setting polymer as it gets hardened and sets during

moulding process and cannot be softened again. The examples are bakelite and melamine-formaldehyde polymers.

- 15.12** (i) The monomer of polyvinyl chloride is $\text{CH}_2=\text{CHCl}$ (vinyl chloride).
 (ii) The monomer of teflon is $\text{CF}_2=\text{CF}_2$ (tetrafluoroethylene).
 (iii) The monomers involved in the formation of bakelite are HCHO (formaldehyde) and $\text{C}_6\text{H}_5\text{OH}$ (phenol).

- 15.14** From the structural point of view, the natural rubber is a linear cis-1,4- polyisoprene. In this polymer the double bonds are located between C_2 and C_3 of isoprene units. This cis-configuration about double bonds do not allow the chains to come closer for effective attraction due to weak intermolecular attractions. Hence, the natural rubber has a coiled structure and shows elasticity.

- 15.16** The monomeric repeat unit of Nylon-6 polymer is:



The monomeric repeat unit of Nylon-6,6 polymer is derived from the two monomers, hexamethylene diamine and adipic acid.

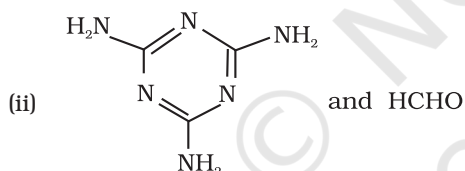


- 15.17** The names and structures of monomers are:

Polymers	Monomer Names	Monomer Structures
(i) Buna-S	1,3-Butadiene Styrene	$\text{CH}_2=\text{CH}-\text{CH}=\text{CH}_2$ $\text{C}_6\text{H}_5\text{CH}=\text{CH}_2$
(ii) Buna-N	1,3- Butadiene Acrylonitrile	$\text{CH}_2=\text{CH}-\text{CH}=\text{CH}_2$ $\text{CH}_2=\text{CH}-\text{CN}$
(iii) Neoprene	Chloroprene	$\begin{array}{c} \text{Cl} \\ \\ \text{CH}_2=\text{C}-\text{CH}=\text{CH}_2 \end{array}$
(iv) Dacron	Ethylene glycol Terephthalic acid	$\text{OHCH}_2-\text{CH}_2\text{OH}$ $\text{COOH}-\text{C}_6\text{H}_4-\text{COOH}$

- 15.18** The monomers forming the polymer are:

- (i) Decandioic acid $\text{HOOC}-(\text{CH}_2)_8-\text{COOH}$ and Hexamethylene diamine $\text{H}_2\text{N}(\text{CH}_2)_6\text{NH}_2$



- 15.19** The following are the equations for the formation of Dacron.

